

Synthesis and chemical investigation of $\text{Sg}(\text{CO})_6^*$

J. Even^{†1}, A. Yakushev², Ch.E. Düllmann^{1,2,3}, H. Haba⁴, M. Asai⁵, T.K. Sato⁵, H. Brand², A. Di Nitto³, R. Eichler^{6,7}, F.L. Fan⁸, W. Hartmann², M. Huang⁴, E. Jäger², D. Kaji⁴, J. Kanaya⁴, Y. Kaneya⁵, J. Khuyagbaatar¹, B. Kindler², J.V. Kratz³, J. Krier², Y. Kudou⁴, N. Kurz², B. Lommel², S. Miyashita^{5,9}, K. Morimoto⁴, K. Morita^{4,10}, M. Murakami^{4,11}, Y. Nagame⁵, H. Nitsche^{12,13}, K. Ooe¹¹, Z. Qin⁸, M. Schädel⁵, J. Steiner², T. Sumita⁴, M. Takeyama⁴, K. Tanaka⁴, A. Toyoshima⁵, K. Tsukada⁵, A. Türler^{6,7}, I. Usoltsev^{6,7}, Y. Wakabayashi⁴, Y. Wang⁸, N. Wiehl^{1,3}, and S. Yamaki^{4,14}

¹Helmholtz-Institut Mainz, Mainz, Germany; ²GSI, Darmstadt, Germany; ³Johannes Gutenberg-Universität, Mainz, Germany; ⁴Nishina Center for Accelerator-Based Science, RIKEN, Wako, Japan; ⁵Advanced Science Research Center, JAEA, Tokai, Japan; ⁶University of Berne, Switzerland.; ⁷PSI, Villigen, Switzerland; ⁸Institute of Modern Physics, Lanzhou, Chinese Academy of Sciences, China; ⁹Hiroshima University, Japan; ¹⁰Kyushu University, Japan; ¹¹Niigata University, Japan; ¹²University of California, Berkeley, CA, U.S.A.; ¹³Lawrence Berkeley National Laboratory, Berkeley, CA, U.S.A; ¹⁴Saitama University, Japan

Gas phase chemical studies of the superheavy elements have been limited to simple inorganic compounds so far [1]. Due to challenging experimental conditions, access to other compound classes was limited. With the combination of physical preseparation and gas-phase chemistry, many limitations could be overcome [2,3]. We succeeded in the synthesis of a carbonyl complex of a superheavy element - seaborgium hexacarbonyl ($\text{Sg}(\text{CO})_6$), at the GAs-filled Recoil Ion Separator GARIS [4]. $\text{Sg}(\text{CO})_6$ has been predicted to be stable [5] and its adsorption behavior on a SiO_2 surface is expected to be very similar to that of $\text{W}(\text{CO})_6$ [6]. Thus, we investigated $\text{Sg}(\text{CO})_6$ along with $\text{W}(\text{CO})_6$. 6-s ^{164}W , and ≈ 10 -s ^{265}Sg were synthesized in the reactions $^{144}\text{Sm}(^{24}\text{Mg},4n)^{164}\text{W}$ and $^{248}\text{Cm}(^{22}\text{Ne},5n)^{265}\text{Sg}$. Evaporation residues (EVRs) were separated from the primary beam and lighter transfer products within GARIS. At the focal plane of GARIS, a recoil transfer chamber (RTC) was installed. The EVRs passed the entrance window of the RTC and were thermalized in a He / CO atmosphere (≈ 600 mbar) in the RTC. The free single ions of W and Sg reacted with CO, forming volatile complexes [7]. The RTC was flushed continuously, transporting volatile compounds through a 10-m long capillary to the Cryo Online Multidetector for Physics and Chemistry of the Transactinides COMPACT [8], a thermochromatography detector array. The chromatography channel is formed by 32 pairs of silicon PIN diodes covered with a SiO_2 surface, kept at temperatures between 22°C and -140°C . Volatile compounds adsorb at a certain temperature on the detector surface. The deposition pattern was compared with Monte Carlo Simulations MCS, which allowed determining the adsorption enthalpy $-\Delta H_{\text{ads}}$. W and Sg were transported

to COMPACT, hence formed volatile compounds with the CO [7]. In total, 15 decay chains assigned to the decay of ^{265}Sg plus three uncorrelated fission event assigned to the decay of ^{261}Rf as a daughter of ^{265}Sg were observed under background-free conditions. The total beam integral was $1.52 \cdot 10^{19}$. Both, the W and the Sg complexes deposited mainly in the last third of the detector (see Fig. 1). The Sg species show the same adsorption behavior as $\text{W}(\text{CO})_6$, which supports the assignment to $\text{Sg}(\text{CO})_6$ [7]. The experimental distributions and the MCS are shown in Fig. 1.

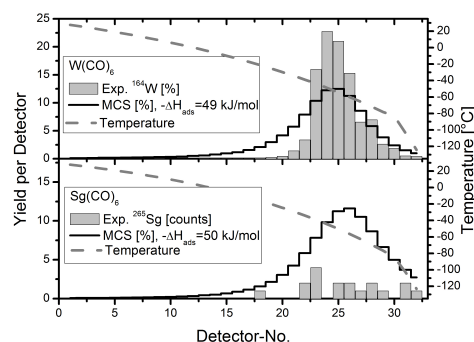


Figure 1: Distribution of ^{164}W and ^{265}Sg in the COMPACT detector array. The bars show the experimental distributions, the solid lines show the results from MCS. (after [7]).

References

- [1] A. Türler, V. Pershina, Chem. Rev. 113, 1237 (2013).
- [2] Ch.E. Düllmann, Nucl. Instr. Meth. A 551, 528 (2005).
- [3] J. Even et al., Inorg. Chem. 51, 6431 (2012).
- [4] K. Morita et al., Nucl. Instr. Meth. B 70, 220 (1992).
- [5] C.S. Nash, B.E. Bursten, J. Am. Chem. Soc. 121, 10830 (1999).
- [6] V. Pershina, J. Anton, J. Chem. Phys. 138, 174301–6 (2013).
- [7] J. Even et al., Science 345, 1491 (2014).
- [8] A. Yakushev et al., Inorg. Chem., 53, 1624 (2014).

* Work supported by HI Mainz, the German Ministry of Research and Education BMBF contract-No. 06MZ7164, the Reimei Research Program (Japan Atomic Energy Agency), the Swiss National Science Foundation, and the U.S. Department of Energy, Office of Science, Chemical Sciences, Geosciences, & Biosciences (CSGB), Division, Heavy Element Chemistry program. This work was performed at the RI Beam Factory operated by RIKEN Nishina Center and CNS, University of Tokyo. We thank the ion source and accelerator staff for providing intense and stable ion beams.

[†] jeven@triumf.ca, present address: TRIUMF, Vancouver, Canada